



Solapur University, Solapur

B. Sc.-III (Chemistry) CREDIT-GRADING SYSTEM Syllabus- 2016-2017

General Structure

Theory Examination:

- Structure of B.Sc. course under faculty of science has total 06 semesters for 3 years.
- B.Sc.-III comprises of total two semesters (Sem-V and Sem-VI).

Each semester will have Five theory papers (one compulsory English and four Chemistry papers) of 70 marks each (University external examination) and 30 marks for each paper (Internal examination)

The duration of each University theory paper examination will be of 2 hr. and 30 min.

Each theory paper has 30 marks for internal examination. There will be 15 marks unit test and 15 marks home assignment

- At the end of academic year i.e. semester - VI the practical examination will be conducted.

The Weightage of practical is of 280 marks for University external practical examination and 120 i.e (30*4) marks for internal practical examination.

There will be Four theory papers in chemistry of 70 marks for each semester. There titles and marks distribution are as under (Excluding English).

Semester	Paper No.	Title of Paper	Lectures/ Practicals per week	Total Marks			Total Credit
				Univ. Exam	Internal Exam	Total	
Semester -V	VII	Physical Chemistry	3	70	30	100	3
	VIII	Inorganic Chemistry	3	70	30	100	3
	IX	Organic Chemistry	3	70	30	100	3
	X	Analytical and Industrial Physical Chemistry	3	70	30	100	3
Semester -VI	XI	Physical Chemistry	3	70	30	100	3
	XII	Inorganic Chemistry	3	70	30	100	3
	XIII	Organic Chemistry	3	70	30	100	3
	XIV	Analytical and Industrial Organic Chemistry	3	70	30	100	3
Theory				560	240	800	24
Practical's			20	280	120	400	20
Total B.Sc.-III				840	360	1200	44

Practical Course

Practical Examination will be held at the end of the year.

A) Distribution of marks :

- **Continuous Internal Assessment for chemistry:**

- 1) Practical paper has $30 \times 4 = 120$ marks for internal examination.
- 2) Practical paper has 280 marks for external university practical examination.
- 3) There will be three practical's, one from each Physical, Inorganic and Organic practical work.
- 4) The mark distribution of 280 marks for external university practical examination is as follows.

Q. 1 Physical Chemistry experiment	: 90 marks
Q. 2 Inorganic Chemistry experiment	: 100 marks
Q. 3 Organic Chemistry experiment	: 90 marks

Total marks: 280 marks

Duration of practical examination is Three days, six and half hours per day

All answer sheets should be collected at the end of examination.

Practical Marks Distribution

- **Physical Chemistry experiment: 90 marks**

- a) Instrumental : 30
- b) Non-instrumental : 40
- c) Journal : 10
- d) Oral : 10

- **Inorganic Chemistry experiment: 100 marks**

- a) Gravimetric analysis : 35
- b) Volumetric analysis : 25
- c) Preparation : 20
- d) Journal : 10
- e) Oral : 10

- **Organic Chemistry experiment: 90 marks**

- a) Organic Mixture Separation and analysis: 35
- b) Volumetric analysis : 25

Or

- c) Preparation : 25
- d) Derivative : 10
- e) Journal : 10
- f) Oral : 10

CHEMISTRY: Syllabus for B.Sc.-III as per CGPA pattern**Theory****N. B.**

- i.) Figures shown in bracket indicates the total number of contact hours required for the respective topics
- ii) The question paper should cover the entire syllabus. Marks allotted should be in proportion to the number of contact hours allotted to respective topics.
- iii) All topics should be dealt with S.I units.
- iv) Use of scientific calculator is allowed.
- v) Industrial tour is prescribed.
- vi) Values required for spectral problems should be provided in the question paper.

SEMESTER –V**PAPER-VII: PHYSICAL CHEMISTRY****Total Credits: 3****Contact hrs: 45****1. Phase Equilibria.****[10]**

- 1.1 Introduction
- 1.2 Gibbs phase rule : Phase rule equation and explanation of terms involved in the equation.
- 1.3 Phase diagram, true and metastable equilibria.
- 1.4 One component systems : (i) Water system (ii) Sulphur system with explanation for polymorphism.
- 1.5 Two component systems : (i) Eutectic system : (Ag - Pb system); Desilverisation of lead, (ii) Formation of compound with congruent melting point (FeCl₃ - H₂O)

2. Electromotive force.**[23]****(Convention : Reduction potentials to be used)**

- 2.1 Introduction
- 2.2 Thermodynamics of electrode potentials, Nernst equation for electrode and cell potentials in terms of activities.
- 2.3 Types of electrodes : Description in terms of construction, representation, half cell reaction and emf equation for,
 - i) Metal - metal ion electrode. ii) Amalgam electrode.
 - iii) Metal - insoluble salt electrode. iv) Gas - electrode.
 - v) Oxidation - Reduction electrode.
- 2.4 i) Reversible and Irreversible cells.
 - ii) Chemical cells without transference.
 - iii) Concentration cells
 - a. Electrode concentration cell
 - I) Reversible to cation
 - II) Reversible to anion
 - b. Electrolyte concentration cells without transference
- 2.5 Equilibrium constant from cell emf, Determination of the thermodynamic parameters

such as ΔG , ΔH and ΔS .

2.6 Applications of emf measurements :

- i) Determination of pH of solution using Hydrogen electrode.
- ii) Solubility and solubility product of sparingly soluble salts (based on concentration cell).

2.7 Numerical problems.

3. Photochemistry.

[12]

3.1 Introduction

3.2 Difference between thermal and photochemical processes.

3.3 Laws of photochemistry : Grotthus - Draper law, Lambert law, Lambert - Beer's law (with derivation), Stark - Einstein law.

3.4 Quantum yield, Reasons for high quantum yield (e.g. $H_2 - Cl_2$) and low quantum yield. (e.g. Decomposition of HI and HBr).

3.5 Photosensitized reactions - Dissociation of H_2 , Photosynthesis.

3.6 Photodimerisation of anthracene.

3.7 Jablonski diagram depicting various processes occurring in the excited state :

Qualitative description of fluorescence and phosphorescence.

3.8 Chemiluminescence.

3.9 Numerical problems.

Reference Books:

1. Physical Chemistry by G. M. Barrow, International student Edition, Mc Graw Hill.
2. University General Chemistry by C.N.R. Rao, Macmillan.
3. Physical Chemistry by, R. A. Alberty, Wiley Eastern Ltd.
4. The Elements of Physical Chemistry by P. W. Atkins, Oxford.
5. Principles of Physical Chemistry by S. H. Maron, C. H. Prutton, 4th Edition.
6. Fundamentals of Photochemistry by K.K. Rohatgi-Mukerjee.
7. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Naginchand and Company, Jalandar.
8. Text Book of Physical Chemistry by S. Glasstone, Macmillan India Ltd.
9. Elements of Physical Chemistry by D. Lewis and S. Glassture (Macmillan).
10. Principles of Physical Chemistry by Maron and Lando (Amerind).
11. An Introduction to Electrochemistry by S. Glasstone.
12. Physical Chemistry by W. J. Moore.
13. Essentials of Physical Chemistry, Bahl and Tuli (S. Chand).

PAPER -VIII: INORGANIC CHEMISTRY**Total Credits: 3**
Contact hrs: 45**1. Metal Ligand Bonding in Transition Metal Complexes [13]****I Crystal Field Theory (CFT).**

- 1) Introduction - What is CFT?
- 2) Basic concept of CFT.
- 3) Formation of complexes with Crystal field splitting of 'd' orbitals
 - I. Shapes of d orbitals and their electron density region
 - II. Formation of octahedral Complex with Crystal field splitting of 'd' orbitals
e.g. High spin and low spin octahedral complexes of Co(III).- $[\text{CoF}_6]^{3-}$, $[\text{Co}(\text{NH}_3)_6]^{3+}$.
 - III. Formation of tetrahedral Complex with Crystal field splitting of 'd' orbitals
e.g. $[\text{CoCl}_4]^{2-}$
 - IV. Formation of square planer Complex with Crystal field splitting of 'd' orbitals
e.g. $[\text{Co}(\text{CN})_4]^{2-}$
- 4) Jahn - Teller distortion.
- 5) Factors affecting the Crystal - field splitting.
- 6) Crystal field stabilization energy (Δ): Calculation for octahedral complexes only.
- 7) Applications and limitations of CFT.

II Molecular Orbital Theory (MOT).

1. Introduction.
2. Basic concept
3. Symmetry classes of atomic orbitals
4. Formation of octahedral complex
 - a) Assumptions
 - b) M.O. energy level diagram for hypothetical octahedral complex.
5. Examples: octahedral complexes with sigma bonding only such as-
e.g. $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$, $[\text{FeF}_6]^{3-}$, $[\text{Fe}(\text{CN})_6]^{3-}$, $[\text{CoF}_6]^{3-}$, $[\text{Co}(\text{NH}_3)_6]^{3+}$, $[\text{Ni}(\text{NH}_3)_6]^{2+}$
6. Applications and limitations of MOT.
7. Comparison between CFT and MOT.

2. Nuclear Chemistry [12]

- 1 Nuclear reaction and energetic of nuclear reactions.
- 2 Classification of nuclear reactions and Types of nuclear reactions-
 - i) Artificial transmutation.
 - ii) Artificial radioactivity.
 - iii) Projectile capture reaction.
 - iv) Projectile capture - particle emission reaction.
 - v) Nuclear fission.
 - vi) Nuclear fusion.
- 3) Use of Uranium, Thorium and Plutonium for:
 - a. Nuclear reactor
 - b. Atomic Bomb.

- 4) Applications of radioisotopes as tracers.
 - i) Chemical investigation - Esterification.
 - ii) Structural determination - Phosphorus pentachloride.
 - iii) Analytical Chemistry - isotopic dilution method for determination of volume of blood.
 - iv) Age determination - Dating by C^{14} .

3. Bioinorganic Chemistry.

[08]

- 1) Essential and trace elements in biological process.
 - i) Essential elements
 - a) Macro / major elements
 - b) Micro/trace/minor elements
 - ii) Non-essential elements
- 2) Metalloporphyrins with special reference to haemoglobin and myoglobin.
 - i) Structure of Haemoglobin (Hb)
 - ii) Structure of Myoglobin (Mb)
 - iii) Function of Haemoglobin (Hb) and Myoglobin (Mb) as Oxygen transport form lungs to tissues
 - iv) Function of Haemoglobin as Carry back CO_2 to lungs
 - v) Co-operativity
 - vi) Oxygen binding curve
 - vii) Difference between Haemoglobin (Hb) and Myoglobin (Mb)
- 3) Role of alkali and alkaline earth metal ions with special reference to Na^+ , K^+ and Ca^{2+} .
 - i) Role of Na^+ and K^+
 - ii) Role of Ca^{2+}

4) Inorganic Polymers.

[08]

- 1) Introduction.
- 2) Basic concept and definition.
 - i) polymer
 - ii) Monomer
 - iii) Polymerization
 - iv) copolymer
 - v) degree of polymerization
- 3) Classification of polymers on the basis of -
 - a) origin
 - b) composition
 - c) properties
 - d) use
- 4) Comparison between organic and inorganic polymers.
- 5) Polymer back bone.
- 6) Homoatomic polymer containing - Phosphorus.
- 7) Heteroatomic polymers –
 - (i) Silicones
 - (ii) Phosphonitrilic compounds

(iii) Fluorocarbons.

5. Nanotechnology

[04]

1. Introduction
2. Properties of nanoparticles
3. Applications of nanoparticles

Reference Books :

1. Concise Inorganic Chemistry (ELBS, 5th Edition) - J. D. Lee.
2. Inorganic Chemistry (ELBS, 3rd Edition) D. F. Shriver, P. W. Atkins, C. H. Lang Ford, Oxford University Press, 2nd Edition.
3. Inorganic Chemistry (Harper International, 3rd edition) J. E. Huheey Harper and Row.
4. Basic Inorganic Chemistry : Cotton and Wilkinson.
5. Advanced Inorganic Chemistry (4th Edn.) Cotton and Wilkinson.
6. Concepts and Models of Inorganic Chemistry : Douglas and Mc. Daniel. 3rd Edition. John Wiley publication.
7. Fundamental concepts of Inorganic Chemistry by E. S. Gilreath.
8. Structural principles in inorganic compounds. W. E. Addison.
9. T. B. of Inorganic analysis - A. I. Vogel.
10. Theoretical principles of Inorganic Chemistry - G. S. Manku.
11. Theoretical Inorganic Chemistry by Day and Selbine.
12. Co-ordination compounds S F A Kettle.
13. Modern Aspects of Inorganic Chemistry. E. Sharpe.
14. New guide to Modern Valence Theory by G. I. Brown.
15. Essentials of Nuclear Chemistry by H. J. Arnikar.
16. Organometallic Chemistry by R. C. Mahotra A. Sing, Wiley Eastern Ltd. New Delhi.
17. Inorganic Chemistry by A. G. Sharpe, Addison - Wisley Longman - Inc.
18. Principles of Inorganic Chemistry by Puri, Sharma and Kalia, Vallabh Publication. Pitampur Delhi.
19. Text book of Inorganic Chemistry by K. N. Upadhyaya Vikas Publishing House - New Delhi.
20. Progress in inorganic polymer by Laport and Leigh.
21. Co-ordination compounds by Baselo and Pearson.
22. Organometallic Chemistry by P. L. Pauson.
23. Advanced inorganic chemistry, Vol. I and II Satyaprakash, G. D. Tuli, S. K. Basu and Madan (S Chand)
24. Selected Topics in inorganic chemistry by W U Malic, G. D. Tuli, R. D. Madan. (S. Chand)
25. Industrial chemistry part I and II by A. K. De
26. Industrial chemistry by B. K. Sharma
27. Nanomaterials and nanostructure, Laura Castlow, April Feter (Dominant Publisher, 2007)
28. Nansocale materials in Chemistry K,J, Kalbunde Wiley Interscience (2001)
29. Introduction to Nansocience and Nanotechnology K.K. Chattopadhyaya, A.N. Banerjee, PHI Learning Pvt. Ltd. New Delhi, 2009.
30. Introduction to Nanotechonology – Charles P. Poole Jr., Frank J Owens, Wiley Student Edition 2008.
31. Nanotechonology – A Future Techonology with Futures BPB Publication New Delhi.

Paper – IX: Organic Chemistry

Total Credits: 3
Contact hrs: 45

1 Spectroscopic Methods.

[20]

1.1. Infrared Spectroscopy :

- 1.1.1 Introduction,
- 1.1.2 Principle of IR spectroscopy,
- 1.1.3 Double beam IR spectrophotometer- Schematic diagram.
- 1.1.4 Fundamental modes of vibration,
- 1.1.5 Types of vibration
- 1.1.6 Hooke's law,
- 1.1.7 factors affecting values of vibrational frequencies,
- 1.1.8 conditions for absorption of radiation and selection rule,
- 1.1.9 fundamental group regions of IR spectrum,
- 1.1.10 Functional group region, Finger print region,
- 1.1.11 characteristic absorption of various functional groups,
- 1.1.12 Applications of IR spectroscopy – Determination of structure, Identification of functional groups spectral problems based on IR

1.2 NMR Spectroscopy.

- 1.2.1 Introduction
- 1.2.2. Proton magnetic resonance (^1H) spectroscopy (PMR).
- 1.2.3 Principles of PMR spectroscopy.
- 1.2.4 Magnetic and non-magnetic nuclei.
- 1.2.5. Theory of PMR spectroscopy - spinning nuclei, magnetic moment and magnetic field, precessional motion of nuclei without mathematical details, nuclear resonance.
- 1.2.6 NMR - Instrument. Schematic diagram.
- 1.2.7. Shielding and deshielding.
- 1.2.8. Chemical shift, measurement of chemical shift, by delta scale and tau scale.
- 1.2.9. TMS as reference. Advantages of TMS.
- 1.2.10. Peak area (integration)
- 1.2.11. Spin - spin splitting ($n + 1$ rule).
- 1.2.12. Definition of coupling constant (J value) of first order coupling.
- 1.2.13. PMR spectra of ethanol, ethyl bromide, acetaldehyde, 1, 1, 2 - tribromoethane, ethyl acetate, acetophenone, benzaldehyde, propanoic acid and benzoic acid
- 1.2.14. Problems pertaining to the structure elucidation of simple organic compounds using PMR spectroscopic data (supporting IR and UV data to be given).

1.3 Mass spectroscopy.

- 1.3.1 Introduction
- 1.3.2 Theory of mass spectroscopy
- 1.3.3 Mass spectrometer - schematic diagram
- 1.3.4 Formation of ions by ionization
- 1.3.5 Types of ions with at least one example.
- 1.3.6. Applications of mass spectroscopy.
 - i) Determination of molecular weight.
 - ii) Determination of molecular formula

2. Stereochemistry.

[07]

- 2.1 Introduction.
 2.2 Baeyer's strain theory.
 2.3 Theory of strainless rings.
 2.4 Conformation and stability of cyclohexane and monosubstituted cyclohexanes , methyl cyclohexane.
 2.5 Locking of conformation in t-butyl cyclohexane.
 2.6 Stereoselective and stereospecific reactions :
 i) Stereochemistry of addition of halogens to alkenes : syn and anti addition.
 Example - Addition of bromine to 2-butene. (mechanism not expected)
 ii) Stereochemistry of elimination reaction : syn and anti elimination Example - Dehydrohalogenation of 1-bromo -1, 2 - diphenylpropane. (Mechanism not expected)

3. Name reactions.

[10]

Mechanism of following reactions :

- 3.1 Stobbe condensation.
 3.2 Oppenauer oxidation.
 3.3 Meerwein Ponndorf Verley reduction.
 3.4 Reformatsky reaction
 3.5 Wagner - Meerwein Rearrangement.
 3.6 Hofmann rearrangement reaction.
 3.7 Wittig reaction.
 3.8 Related problems.

4. Organic synthesis via Enolates

[08]

- 4.1 Introduction - Reactive methylene group.
 4.2 Ethyl acetoacetate - synthesis by Claisen condensation, acidity of methylene hydrogen (salt formation), Keto-enol tautomerism, synthetic applications - Synthesis of alkyl and dialkyl derivatives, monobasic, dibasic and α - β - unsaturated acid, heterocyclic compound.
 4.3 Diethyl malonate - Synthesis, acidity of methylene hydrogen (salt formation). Synthetic applications - Synthesis of alkyl and dialkyl derivatives, monobasic , dibasic acid, α - β - unsaturated acid, α -amino acid and heterocyclic compound.

Reference Books :

- 1) Organic Chemistry - Cram D. J. and Hammond G.S. McGraw Hill book Company New York.
- 2) Organic Chemistry - Finar I. L. The English Language Book Society, London.
- 3) A Guide Book to mechanism in Organic Chemistry - Peter Sykes Longman Green and Co. Ltd. London 6th Edition.
- 4) Organic Chemistry - R. T. Morrison and R. N. Boyd Prentice Hall of India private limited New Delhi. 6th Edition.
- 5) Text book of organic Chemistry - Furguson L. N. D. Van Nostrand Company Indian Edition, Affiliated East West press private Ltd. New Delhi.
- 6) Organic Chemistry Vol. I, II and III - S. M. Mukharji, S. P. Singh, R. P. Kapoor Wiley Estern, Limited, New Delhi.

- 7) A text book of organic Chemistry - K. S. Tewari, S. N. Mehrotra, N. K. Vishnoi
Vikas Publishing House Private Ltd. New Delhi.
- 8) A text book of Organic Chemistry - Arun Bahl and B. S. Bahl S. Chand and
Company Ltd. 6th Edition.
- 9) Heterocyclic Chemistry Synthesis, Reactions and Mechanism - Raj K. Bansal
Wiley Easter Ltd. New Delhi.
- 10) Reaction Mechanism and reagents in Organic Chemistry - G. R. Chatwal Himalaya
Publishing House New Delhi.
- 11) Stereochemistry conformation and mechanism - P. S. Kalsi, New Age International
Publishers 4th Edition.
- 12) Organic Chemistry Volume I and II - I. L. Finar ELBS with Longman 6th Edition.
- 13) Organic Chemistry Volume I and II - William Kemp ELBS with Macmillan 3rd Edition.
- 14) Advanced Organic Chemistry - Jerry March Wiley Eastern Ltd.
- 15) Spectroscopy of Organic compounds - P. S. Kalsi.
- 16) Organic Chemistry - Fieser and Fieser.
- 17) Principles of Organic Chemistry - English and Cassidy.
- 18) Elementary Organic Absorption Spectroscopy - Y. R. Sharma.
- 19) Spectroscopy - V. M. Parikh.
- 20) Stereochemistry of Carbon Chemistry - Eliel.
- 21) Principles of Organic Chemistry - M. K. Jain.
- 22) Organic Chemistry by Clayden, Greeves, Warren and Wothers Oxford press.
- 23) Organic Chemistry - A Comprehensive degree text and source book by Hanes
Baeyers and Wolfgang Walter Albion Chemical Science Series.
- 24) Reactions, Rearrangements and reagents - S.N.Sanyal Bharati Bhawan publishers and
Distributors Patna.

PAPER-X : ANALYTICAL AND INDUSTRIAL PHYSICAL CHEMISTRY

Total Credits: 3
Contact hrs: 45

1. Colorimetry. [08]

1.1 Introduction

1.2 General discussion of theory of colorimetry : Lambert law, Beer's law (Derivation not expected), Terms used in Colorimetry, Application of Beer's law, Deviation from Beer's law.

1.3 Classification of methods of 'colour' measurement or comparison, Photoelectric photometer method - single cell photo-electric colorimeter.

2. Potentiometry [10]

2.1 Introduction.

2.2 Detail study of calomel, quinhydrone and glass electrodes and their use in determination of pH. .

2.3 Potentiometric titrations : Classical and analytical methods for locating end points,

i) Acid - Base titrations.

ii) Redox - titrations.

iii) Precipitation titrations.

2.4 Advantages of potentiometric titrations,

2.5 Basic circuit of direct reading potentiometer.

3 Electroplating [08]

3.1 Introduction.

3.2 Electrolysis, Faraday's laws, Cathode current efficiency.

3.3 Basic principles of electroplating, cleaning of articles.

3.4 Electroplating of Nickel and Chromium.

3.5 Anodising.

4 Flame photometry [09]

4.1 General principles.

4.2 Instrumentation : Block diagram, Burners: Total consumption burner, premix or laminar-flow burner, Lundergraph burner,

Mirrors, Slits, Monochromators, Filters and Detectors.

4.3 Applications in qualitative and quantitative analysis.

4.4 Limitations of flame photometry.

5. Conductometry: [10]

5.1 Measurement of conductance by Wheatstone bridge, Basic circuit of D.C. Wheatstone bridge, use of alternating current,

conductivity water, Different types of conductivity cells, cell constant and its determination.

Experimental determination of specific, equivalent and molecular conductance's.

5.2 Conductometric acid-base titrations

i. Strong acid against strong base

ii. Strong acid against weak base

iii. Weak acid against strong base.

iv. Weak acid against weak base.

5.3 Advantages of conductometric titrations

Reference Books :

1. Text book of Quantitative Inorganic Analysis - By A. I. Vogel (ELBS and Longman 3rd Edition).
2. Instrumental methods of Chemical analysis by Willard, Merit and Dean.
3. Instrumental methods of Chemical analysis by Chatwal and Anand (Himalaya Publication).
4. Principles of electroplating and eletroforming by Blum and Hogaboom, Mac Graw - Hill Book Co. 3rd Edn.
5. Vogel's text book of Quantitative Inorganic Analysis by Basssett and Denny etc. ELBS and Longman 4th Edition.
6. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Naginchand and Company, Jalandar.
7. Text Book of Physical Chemistry by S. Glasstone, Macmillan India Ltd.
8. Elements of Physical Chemistry by D. Lewis and S. Glasstone (Macmillan).
9. Principles of Physical Chemistry by Maron and Lando (Amerind).
10. An Introduction to Electrochemistry by S. Glasstone.
11. Physical Chemistry by W. J. Moore.
12. Essentials of Physical Chemistry, Bahl and Tuli (S. Chand).

SEMESTER-VI**PAPER- XI: PHYSICAL CHEMISTRY****Total Credits: 3
Contact hrs: 45****1. Spectroscopy. [10]**

1.1 Introduction

1.2 Electromagnetic radiation.

1.3 Electromagnetic spectrum, Energy level diagram.

1.4 Rotational spectra of diatomic molecules : Rigid rotor model; moment of inertia (derivation not expected); energy levels of rigid rotor, selection rule; spectral intensity; distribution using population distribution (Maxwell - Boltzman distribution), determination of bond length; isotope effect. Interaction of radiation with rotating molecule.

1.5 Vibrational spectra of diatomic molecules : Simple Harmonic oscillator model, Vibrational energies of diatomic molecules, Determination of force constant, overtones. Interaction of radiation with vibrating molecules.

1.6 Numerical problems.

2. Solutions. [09]

2.1 Introduction

2.2 Ideal solutions, Raoult's law, vapour pressure of ideal and non ideal solutions of miscible liquids.

2.3 Vapour pressure and boiling point diagrams of miscible liquids.

Type I : Systems with intermediate total vapour pressure.

(i.e. System in which B.P. increases regularly - Zeotropic)

Type II : Systems with a maximum in the total vapour pressure.

(i.e. System with a B.P. minimum - Azeotropic)

Type III : Systems with a minimum in the total vapour pressure.

(i.e. System with a B.P. Maximum - Azeotropic)

Distillation of miscible liquid pairs.

2.4 Solubility of partially miscible liquids.

(i) Maximum solution temperature type : Phenol - water system.

(ii) Minimum solution temperature type : Triethyl amine - water system.

(iii) Maximum and minimum solution temperature type : Nicotine - water system.

3. Thermodynamics. [13]

3.1 Introduction

3.2 Free energy : Gibbs function (G) and Helmholtz function (A), Criteria for thermodynamic equilibrium and spontaneity.

3.3 Relation between ΔG and ΔH : Gibbs Helmholtz equation.

3.4 Phase equilibria : Clapeyron - Clausius equation.

3.5 Thermodynamic derivation of law of mass action, van't - Hoff isotherm and isochore.

3.6 Fugacity and activity concepts.

3.7 Numerical problems.

4. Chemical Kinetics**[13]**

4.1 Introduction, simultaneous reactions such as opposing reactions, side reactions, consecutive reactions and chain reactions

[Derivations of rate Equations for these reactions are not expected.]

4.2 Effect of temperature on the rate of reaction.

1. Temperature coefficient
2. Arrhenius equation
3. Energy of activation

4.3 Theories of reaction rate :

1. Collision theory and
2. Transition state theory

4.4 Third order reaction with equal concentration of all reactants, their characteristics and examples

4.5 Numerical problems.

Reference Books :

1. Principles of Physical Chemistry by Maron and Pruton 4th edition.
2. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Naginchand and Company, Jalandar.
3. Text Book of Physical Chemistry by S. Glasstone, Macmillan India Ltd.
4. Elements of Physical Chemistry by D. Lewis and S. Glasstone (Macmillan).
5. Principles of Physical Chemistry by Maron and Lando (Amerind).
6. Thermodynamics for chemists by S Glasstone.
7. Physical Chemistry by W. J. Moore.
8. Essentials of Physical Chemistry, Bahl and Tuli (S. Chand).
9. Basic Chemical Thermodynamics by V V Rao (Macmillan)
10. An introduction to chemical thermodynamics by R. R. Mishra and R. P. Rastogi.
11. Fundamentals of molecular spectroscopy by C. N. Banwell and McCash- Tata McGraw Hill

PAPER-XII: INORGANIC CHEMISTRY**Total Credits: 3****Contact hrs: 45****1) Study of F-block Elements [11]**

1.1 Lanthanides :-

- I) Introduction
- II) Electronic configuration
- III) Occurrence
- IV) Separation of Lanthanides
 - i) Bulk separation methods
 - ii) Individual separation of lanthanides- Mention names of methods only; In detail Ion exchange method.

1.2 Actinides :-

- I) Introduction
- II) Electronic configuration
- III) General Methods of preparation –
 - a. Neutron-capture followed by β -decay
 - b. Accelerated projectile bombardment method
 - c. Heavy-Ion Bombardment method

1.3 IUPAC Nomenclature of the super Heavy Elements with atomic numbers (Z) greater than 100.

2) Metals and Semiconductors. [11]

2.1 Introduction.

2.2 Properties of metallic solids.

2.3 Theories of bonding in metal.

- a) Free electron theory.
- b) Molecular orbital theory (Band theory).

2.4 Classification of solids as conductor, insulators and semiconductors on the basis of band theory.

2.5 Semiconductors:

- a) Types of semiconductors - intrinsic and extrinsic semiconductors.
- b) Applications of semiconductors.

2.6 Superconductors :

- a) Ceramic superconductors - Preparation and structures of mixed oxide $\text{YBa}_2\text{Cu}_3\text{O}_{7-x}$
- b) Applications of superconductors.

3) Structural Chemistry. [11]

3.1 Structural study of following compounds.

- i) Diborane.
- ii) Borazine.
- iii) Xenon compounds $\rightarrow \text{XeF}_2, \text{XeF}_6, \text{XeO}_4$ (w.r.t. VBT only.)

3.2 Structural study of Oxides Of Sulphur and Phosphorous:

- I Oxides of Sulphur : SO_2 , and SO_3 ;
- II Oxides of Phosphorous : P_4O_6 and P_4O_{10}

4) Corrosion and Passivity.**[07]**

4.1 Corrosion :-

- A) Introduction, with types of corrosion.
- B) Electrochemical theory of corrosion.
- C) Factors affecting the corrosion.
 - i) Position of metal in emf series.
 - ii) Purity of metal.
 - iii) Effect of moisture.
 - iv) Effect of oxygen.
 - v) Hydrogen over voltage.
- D) Methods of protection of metals from corrosion.

4.2 Passivity :-

- A) Definition.
- B) Types of passivity.
- C) Oxide film theory.
- D) Application of passivity.

5. Organometallic Chemistry.**[05]**

5.1 Introduction - Definition,

5.2 Nomenclature of organometallic compounds.

5.3 Synthesis and structural study of alkyl and aryl compounds of Li, Be and Al.

5.4 Mononuclear carbonyl and nature of bonding in simple metal carbonyls.

Reference Books :

1. Concise Inorganic Chemistry (ELBS, 5th Edition) - J. D. Lee.
2. Inorganic Chemistry (ELBS, 3rd Edition) D. F. Shriver, P. W. Atkins, C. H. Lang Ford, Oxford University Press, 2nd Edition.
3. Inorganic Chemistry (Harper International, 3rd edition) J. E. Huheey Harper and Row.
4. Basic Inorganic Chemistry : Cotton and Wilkinson.
5. Advanced Inorganic Chemistry (4th Edn.) Cotton and Wilkinson.
6. Concepts and Models of Inorganic Chemistry : Douglas and Mc. Daniel. 3rd Edition. John Wiley publication.
7. Fundamental concepts of Inorganic Chemistry by E. S. Gilreath.
8. Structural principles in inorganic compounds. W. E. Addison.
9. T. B. of Inorganic analysis - A. I. Vogel.
10. Theoretical principles of Inorganic Chemistry - G. S. Manku.
11. Theoretical Inorganic Chemistry by Day and Selbine.
12. Co-ordination compounds S F A Kettle.
13. Modern Aspects of Inorganic Chemistry. E. Sharpe.
14. New guide to Modern Valence Theory by G. I. Brown.
15. Essentials of Nuclear Chemistry by H. J. Arnikar.
16. Organometallic Chemistry by R. C. Mahrotra A. Sing, Wiley Eastern Ltd. New Delhi.
17. Inorganic Chemistry by A. G. Sharpe, Addison - Wisley Longman - Inc.
18. Principles of Inorganic Chemistry by Puri, Sharma and Kalia, Vallabh Publication. Pitampur Delhi.
19. Text book of Inorganic Chemistry by K. N. Upadhyaya Vikas Publishing House – New Delhi.

20. Progress in inorganic polymer by Laport and Leigh.
21. Co-ordination compounds by Baselo and Pearson.
22. Organometallic Chemistry by P. L. Pauson.
23. Advanced inorganic chemistry, Vol. I and II Satyaprakash, G. D. Tuli, S. K. Basu and Madan
24. Selected Topics in inorganic chemistry by W U Malic, G. D. Tuli, R. D. Madan. (S. Chand)
25. Industial chemistry part I and II by A. K. De
26. Industrial chemistry by B. K. Sharma

Paper - XIII : Organic Chemistry**Total Credits: 3****Contact hrs: 45****1 Heterocyclic compounds****[09]**

1.1 Introduction and classification.

1.2 Pyrrole.

1.2.1 Methods of synthesis :

i) From acetylene.

ii) From furan.

iii) From succinamide.

1.2.2 Physical properties.

1.2.3 Reactivity of pyrrole :

i) Basic character.

ii) Acidic character.

iii) Electrophilic substitution with general mechanism.

1.2.4 Chemical reactions :

i) Reduction.

ii) Oxidation.

iii) Nitration, sulphonation and halogenation.

iv) Friedel Craft's reaction.

v) Coupling reaction.

1.3 Pyridine.

1.3.1 Methods of synthesis.

i) From acetylene and hydrogen cyanide.

ii) From piperidine.

1.3.2 Physical properties.

1.3.3 Chemical reactions

i) Basic character

ii) Electrophilic substitution(nitration, sulphonation and bromination) reactions

iii) Nucleophilic substitution - General mechanism, Reactions with sodamide, sodium hydroxide and n-Butyl lithium.

1.4 Quinoline

1.4.1 Synthesis - Skraup's synthesis

1.4.2 Physical properties.

1.4.3 Reactions of quinoline :

i) Electrophilic substitution reactions - Nitration and sulphonation.

ii) Nucleophilic substitution reactions - Reactions with sodamide, alkylation and arylation.

iii) Reduction.

2. Carbohydrates**[11]**

2.1 Introduction.

2.2 Classification and nomenclature.

2.3 Monosaccharide D-glucose - Open chain structure.

2.4 Chain lengthening of Aldoses - Kiliani synthesis.

2.5 Chain shortening of Aldoses - Weerman's reaction.

2.6 Interconversion of glucose and fructose.

2.7 Configuration of D-glucose from D-arabinose.

- 2.8 Objections against open chain structure of D-glucose.
 2.9 Muta-rotation with mechanism.
 2.10 Ring structure of D-glucose - Determination of size of ring by,
 i) Methylation method.
 ii) Periodic acid treatment method.
 iv) X - ray analysis.
 2.11 Disaccharides - Introduction, sucrose and lactose - Sources, structural formulae and uses.
 2.12 Polysaccharides - Introduction starch, - Sources, structural formulae and uses.

3. Vitamins and Hormones

[07]

- 3.1 General idea of vitamins, structure and synthesis of vitamin A
 3.2 General idea of hormones, structure and synthesis of Adrenaline and Thyroxin

4. Pharmaceuticals

[07]

- 4.1 Introduction.
 4.2 Qualities of ideal drug.
 4.3 Methods of classification of drugs - Classification based on the therapeutical action.
 4.4 Brief idea of pencillin-G (constitution, synthesis not expected)
 4.5 Synthesis and uses of the following drugs :
 i) Antimalerials - Paludrin.
 ii) Antituberculars - Isoniazide and Ethambutol.
 iii) C. N. S. drugs - Phenobarbitone.
 iv) Antidiabetics - Tolbutamide.
 v) Antiinflammatory drugs - Ibuprofen.
 vi) Antibiotic - Chloromycetin.

5 Synthetic dyes.

[07]

- 5.1 Introduction, Qualities of good dye.
 5.2. Classification based on constitution and methods of applications.
 5.3 Witt's theory - Colour and constitution.
 5.4 Synthesis of Orange IV, Malechite green, phenolphthalein.

6 Agrochemicals.

[04]

- 6.1 General idea of agrochemicals including pyrethroides.
 6.2 Synthesis and uses of the following agrochemicals :
 i) Indole-3-acetic acid.
 ii) Monocrotophos.
 iii) Methoxychlor.
 iv) Ethophan.
 v) Carbaryl.

Reference Books :

- 1) Organic Chemistry - Cram D. J. and Hammond G.S. McGraw Hill book Company New York.
- 2) Organic Chemistry - Finar I. L. The English Language Book Society, London.
- 3) A Guide Book to mechanism in Organic Chemistry - Peter Sykes Longman Green and Co. Ltd. London 6th Edition.
- 4) Organic Chemistry - R. T. Morrison and R. N. Boyd Prentice Hall of India private limited New Delhi. 6th Edition.
- 5) Text book of organic Chemistry - Furguson L. N. D. Van Nostrand Company Indian Edition, Affiliated East West press private Ltd. New Delhi.
- 6) Organic Chemistry Vol. I, II and III - S. M. Mukharji, S. P. Singh, R. P. Kapoor Wiley Estern, Limited, New Delhi.
- 7) A text book of organic Chemistry - K. S. Tewari, S. N. Mehrotra, N. K. Vishnoi Vikas Publishing House Private Ltd. New Delhi.
- 8) A text book of Organic Chemistry - Arun Bahl and B. S. Bahl S. Chand and Company Ltd. 6th Edition.
- 9) Heterocyclic Chemistry Synthesis, Reactions and Mechanism - Raj K. Bansal Wiley Easter Ltd. New Delhi.
- 10) Reaction Mechanism and reagents in Organic Chemistry - G. R. Chatwal Himalaya Publishing House New Delhi.
- 11) Organic Chemistry Volume I and II - I. L. Finar ELBS with Longman 6th Edition.
- 12) Organic Chemistry Volume I and II - William Kemp ELBS with Macmillion 3rd Edition.
- 13) Advanced Organic Chemistry - Jerry March Wiley Eastern Ltd.
- 14) Spectroscopy of Organic compounds - P. S. Kalsi.
- 15) Organic Chemistry - Fieser and Fieser.
- 16) Principles of Organic Chemistry - English and Cassidy.
- 17) Chemicals for crop improvement and pest management - Green, Hartly and West.
- 18) Chemistry of pesticides - K. H. Buchel (T. W.).
- 19) Medical Chemistry - Burger.
- 20) Principles of Organic Chemistry - M. K. Jain.
- 21) Organic Chemistry by Clayden, Greeves, Warren and Wothers Oxford press.
- 22) Organic Chemistry - A Comprehensive degree text and source book by Hanes Baeyers and Wolfgang Walter Albion Chemical Science Series.
- 23) Reactions, Rearrangements and reagents - S.N.Sanyl Bharati Bhawan publishers and Distributors Patna.
- 24) Synthetic Organic Chemistry-Kamlesh Bansal.
- 25) Synthetic Organic Chemistry-Gurudeep Chatwal.

Paper-XIV: Analytical and Industrial Organic Chemistry**Total Credits: 3****Contact hrs: 45****1. Soaps and Detergents. [08]**

1.1 Soap

- i) Raw materials.
- ii) Types of soaps.
- iii) Manufacture of soap - Hot process.
- iv) Cleansing action of soaps.

1.2 Detergents

- i) Raw materials.
- ii) Types of detergents - Cationic, anionic, amphoteric, neutral detergents.
- iii) Preparation of teepol and deriphat.

1.3 Comparison between soaps and detergents.

2. Synthetic polymers. [08]

2.1 Introduction.

2.2 Classification :

- i) According to origin, composition, method of preparation and general physical properties.

- ii) Classification based upon structure.

2.3 Process of addition polymerisation - free radical polymerisation of alkenes and Dienes.

2.4 Ionic polymerisation.

2.5 Ziegler - Natta polymerisation.

2.6 Methods of preparation and uses of :

- i) Polythene. ii) Polystyrene
- iii) PVC. iv) Phenol formaldehyde resin.
- v) Urea formaldehyde resin vi) Poly urethane

2.7 Natural rubber : General idea and vulcanisation.

2.8 Synthetic rubbers : Synthesis and uses of -

- i) Polychloroprene, ii) Buna rubber - Buna N and Buna S.

3. Sugar and Alcohol Industry [09]

3.1 Manufacture of raw cane sugar.

3.2 Refining of raw sugar.

3.3 White sugar.

3.4 By-products of sugar industry.

3.4.1 Manufacture of ethyl alcohol from molasses

3.4.2 Rectified spirit, Denatured spirit absolute alcohol and power alcohol.

3.4.3 By-products of alcohol industry.

4. Textile chemistry [07]

4.1 Introduction, classification of fibers.

4.2 Sizing: object of sizing, sizing ingredients and their functions.

4.3 General idea of processes like singeing, desizing, scouring.

- 4.4 Bleaching: i) Brief study of the outline of the process of bleaching cotton and synthetic material.
- 4.5 Dyeing : Study of dyeing of cellulosic material and synthetic fibers with dyes like direct, vat, reactive and disperse dyes.

5. Green Chemistry.

[04]

- 5.1 Introduction - Twelve principles of green chemistry.
- 5.2 Zeolites - Friedel Craft's alkylation and acylation, oxidation of benzene to phenol and benzoquinone, Reduction of benzoquinone to hydroquinone.
- 5.3 Biocatalytic reaction - Hydroxylation and oxidation using enzymes.
- 5.4 Introduction to microwave assisted reactions.

6 Chromatography.

[09]

- 6.1 Introduction.
- 6.2 General principles.
- 6.3 Classification.
- 6.4 Study of following chromatographic techniques with reference to principle, methodology and applications.
- i) Paper chromatography.
 - ii) Column chromatography.
 - iii) Thin layer chromatography.
 - iv) Gas chromatography

Reference) Books :

1. Basic Concepts of Analytical Chemistry - S. M. Khopkar, Wiley Eastern Ltd., Bombay.
 2. Industrial Chemistry - R. K. Das, Asia Publishing, Mumbai.
 3. Text Book of Quantitative Organic Analysis - A. I. Vogel, Pearson Edn. Delhi.
 4. Quantitative Organic Chemistry - A. I. Vogel, Pearson Edn. Delhi.
 5. Hand Book of Organic Analysis - H. T. Clarke, Arnold Heinemann Pub. Delhi.
 6. Advanced Organic Chemistry - B. S. Bahl and Arun Bahl, S. Chand Comp. Delhi.
 7. Riegel's Handbook of Industrial Chemistry - J. A. Kent, Van. Nostrand, London.
 8. Chemical Process Industries - Shreve and Brinic - Ostin, Magraw Hill, New York.
 9. Analytical Chemistry- Walton.
 10. Biotechnology and Applied Microbiology - Alani and Moo-Young.
 11. Immobilize Biocatalysis - Joy Wleser.
 12. Introduction to Polymer Chemistry - Raymond B. Seymour.
 13. Polymer Science - V. R. Gowarikar, N. V. Viswanathan and Jayadev Sreedhar Willey Eastern Limited.
 14. Advances in Green Chemistry : Chemical synthesis using MW-irradiation by R. S. Varma.
 15. Green Chemistry : Environment Friendly alternatives - Rashmi Sanghi and M. M. Srivastava (Eds) (c) 2003 Narosa Publishing House, New Delhi, India.
 16. Textile science - J. T. Marsh
 17. A book of textile chemistry - A. J Hall, (Butterworths group)
 18. Sizing - D..B. Ajagaonkar
 19. Bleaching - Dr. V. A. Shenai
 20. Dyeing - Dr. V. A. Shenai
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Practicals

- N. B. i. Use of Analytical / Chainometric / Electronic / Single pan balance is allowed.
 ii. Use of Scientific calculator is allowed.

Physical Chemistry

I) Non instrumental Experiments(Any Five) :

1. To determine the equilibrium constant of the reaction, $KI + I_2 = KI_3$ by the distribution method.
2. To determine the partition coefficient of CH_3COOH between H_2O and CCl_4 .
3. Critical Solution Temperature.
To determine the CST for phenol – water system.
4. The study of energy of activation of first order reaction i.e. hydrolysis of methyl acetate in presence of 0.5 N HCl.
5. The study of energy of activation of first order reaction i.e. hydrolysis of methyl acetate in presence of 0.5 N H_2SO_4 .
6. The study of energy of activation of second order reaction i.e. reaction between $K_2S_2O_8$ and KI (Equal concentrations).
7. The study of energy of activation of second order reaction i.e. reaction between $K_2S_2O_8$ and KI (Unequal concentrations).
8. To study the hydrolysis of methyl acetate by using its two concentrations in presence of 0.5 N HCl and hence find velocity constant of the reaction.
9. To study the effect of addition of electrolyte (KCl) on the reaction between $K_2S_2O_8$ and KI (Equal concentrations).

II. Instrumental experiments

A. Potentiometry (Any Three).

1. Titration of strong acid with strong alkali.
2. Preparation of buffer solution and determination of their pH (Any five buffer solutions), - Theoretical calculation of pH values by using Henderson's equation.
3. Determination of standard electrode potential of Zn/Zn^{++} , Cu/Cu^{++} , Ag/Ag^+ (Any two).
4. Determination of solubility and solubility product of AgCl.
5. Titration of ferrous ammonium sulphate using $K_2Cr_2O_7$ solution and to calculate redox potential of Fe^{++} , Fe^{+++} system

B. Conductometry (any three).

1. Titration of weak acid with strong alkali.
2. Titration of a mixture of weak acid and strong acid with strong alkali.
3. To study the effect of substituent on dissociation constant of weak acid with respect to acetic acid and monochloroacetic acid (cell constant to be given).
4. To determine the velocity constant of hydrolysis of ethyl acetate by NaOH solution by conductometric method.

C. Refractometry.

1. To determine the percentage composition of unknown mixture by (i) graphical method and (ii) by composition law (Densities of pure liquids A & B be given).
2. To determine the molar refractivity of methyl acetate, ethyl acetate, n-hexane and

carbon tetrachloride and calculate the refraction equivalents of C, H and Cl atoms.

D. Colorimetry (Any Two).

1. To verify Lambert - Beer's law using CuSO_4 solution.
2. To estimate of Fe^{+++} ions by thiocyanate method.
3. To estimate Fe^{+++} ions using salicylic acid by colorimetric titration.

E. pH - metry (Any One).

1. To determine the dissociation constant of monobasic acid (Acetic acid).
2. To determine the dissociation constant of dibasic acid (Malonic acid).

Reference Books :

1. Findlay's Practical Physical Chemistry (Longman)
2. Advanced Practical Physical Chemistry by J. B. Yadav, Goel publishing house.
3. Practical Physical Chemistry by B. D. Khosla, V. C. Garg (R. Chand and Co.)
4. Systematic experimental Physical Chemistry by Rajbhoj, Chandekar (Anjali Publicaiton)
5. Practical Physical Chemistry : Nandkumari, Kothari and Lavande.
6. Practical Physical Chemistry by Gurtu (S. Chand).

Inorganic Chemistry

I. Gravimetric Estimations (G).

N. B. : Any two experiments from G1 to G3 and any two experiments from G4 to G7

G1. Gravimetric estimation of iron as ferric oxide from the given solution containing ferrous ammonium sulphate, copper sulphate and free sulphuric acid.

G2. Gravimetric estimation of zinc as zinc pyrophosphate from the given solution containing zinc sulphate, ferrous ammonium sulphate and free sulphuric acid.

G3. Gravimetric estimation of barium as barium sulphate from the given solution containing barium chloride, ferric chloride and free hydrochloric acid.

G4. Gravimetric estimation of manganese as manganese ammonium phosphate from the given solution containing manganese sulphate, copper sulphate and free sulphuric acid.

G5. Gravimetric estimation of barium as barium chromate from the given solution containing barium chloride, ferric chloride and free hydrochloric acid.

G6. Gravimetric estimation of Aluminium as Aluminium oxinate i.e.

tris (8-hydroxyquinolino) aluminate (III) from a given solution containing potash alum, copper sulphate and free sulphuric acid.

G7. Gravimetric estimation of nickel as bis (dimethylglyoximato) nickel (II) from the given solution containing nickel sulphate, ferrous ammonium sulphate and free sulphuric acid.

[For the gravimetric experiments, stock solution should be given in the range of 10 to 15 cm and asked to dilute to 100 cm (or the stock solution should be given in the range of 20 to 30 cm and asked to dilute to 250 cm). Use 50 cm of this diluted solution for estimation.]

II. Inorganic Preparations (P) : (any five).

N. B. - Calculations of % yield is expected.

P1. Preparation of sodium cuprous thiosulphate.

P2. Preparation of potassium trioxalato ferrate (III).

P3. Preparation of potassium trioxalatoaluminate (III).

P4. Preparation of tris (ethylene diamine)nickel (II) thiosulphate.

P5. Preparation of sodium hexanitrocobaltate (III).

P6. Preparation of ammonium diamminetetra thiocyanatochromate(III) (Reineck's salt).

P7. Preparation of chloropentaamminecobalt (III) chloride.

P8. Preparation of hexamminenickel (II) chloride.

P9. Preparation of tris(thiourea)cuprous(I) sulphate.

III) Titrimetric Estimations :

A) Percentage Purity (any three)

V1. Determination of percentage purity of ferrous ammonium sulphate.

V2. Determination of percentage purity of tetramminecopper (II) sulphate.

V3. Determination of percentage purity of potassium trioxalatoaluminate(III).

V4. Determination of percentage purity of potassium trioxalato ferrate (III).

B) Analysis of Commercial Sample (any three).

V5. Determination of percentage of magnesium in the given sample of talcum powder.

V6. Determination of amount of aluminium in the given solution of potash alum.

V7. Determination of titrable acidity in the given sample of milk or lassi.

V8. Determination of Chemical Oxygen Demand of the given sample of industrial effluent by dichromate method.

V9. Determination of percentage purity of boric acid using supplied sodium hydroxide (Standard succinic or oxalic acid solution to be prepared to standardise the given sodium hydroxide solution.)

C) Ion exchange method (any two).

V10. Determination of amount of sodium present in the given solution of common salt using cation exchange resin (By Acid Base titration).

V11. Determination of amount of magnesium in the given solution containing (Mg^{++} and Zn^{++}) using anion exchange resin and standard solution of EDTA.

V12. Determination of amount of zinc in the given solution containing (Mg^{++} and Zn^{++}) using anion exchange resin and standard solution of EDTA.

Reference Books:

1. A text book of quantitative Inorganic Analysis - A. I. Vogel.
2. Text book of Quantitative Inorganic Analysis - Kolthoff and Sandell.
3. Experimental Inorganic Chemistry - Palmer W. G.
4. Advanced Practical Inorganic Chemistry - Adams and Raynor.
5. Handbook of Preparation Inorganic Chemistry. Vol. 1 and 11 - Brauer.
6. Manual in Dairy Chemistry - I.C.A.R. Sub-Committee on Dairy Education.
7. Chemical methods for environmental analysis - R. Ramesh and M. Anbu.

Organic Chemistry

I) Qualitative analysis

Separation of binary mixture and Identification of its components. 5g of mixture is to be given for separation.

At least **08 mixtures** are to be separated.

Nature 1) Solid - Solid : 4 mixtures

2) Solid - Liquid : 2 mixtures

3) Liquid - Liquid : 2 mixtures

1) Solid - Solid Mixtures :

One mixture from each the following types should be given :

i) Acid + Phenol ii) Acid + Base

iii) Acid + Neutral iv) Phenol + Base

v) Phenol + Neutral vi) Base + Neutral

2) Solid - Liquid Mixtures

One mixtures of type Neutral + Neutral and One mixture of type Acid + Neutral should be given.

3) Liquid - Liquid Mixtures

One mixture of type Neutral + Neutral and One mixture of type Base + Neutral should be given.

Following compounds should be used for preparation of mixtures:

Acids : Benzoic acid, Phthalic acid, Salicylic acid, Cinnamic acid, Aspirin, Oxalic acid.

Phenols: α -naphthol, β -naphthol

Bases : o -nitroaniline, m-nitroaniline, p-nitroaniline, aniline, o-toluidine and N, N-dimethyl aniline.

Neutrals : Naphthalene, acetanilide, m-dinitrobenzene, chloroform, carbon tetrachloride, acetone, nitrobenzene, ethyl acetate, ethyl benzoate, acetophenone, bromobenzene, urea and thiourea.

II) Quantitative analysis : (Any four)

Organic estimations :

1) Estimation of sucrose

2) Estimation of nitro group

3) Saponification value of oil.

4) Estimation of formaldehyde from given formalin solution.

5) Estimation of acid and ester present in the given mixture of acid and ester.

6) Estimation of acid and amide from the mixture of acid and amide.

III) Organic Preparations : with following – (any four)

N.B.: a) Calculation of percentage practical yield.

b) Recrystallisation of crude product and its melting point.

c) The purity of the product may be confirmed by TLC.

1) Preparation of m-nitroaniline from m-dinitrobenzene.

2) Preparation of aspirin from salicylic acid .

3) Preparation of nerolin from β -naphthol

4) Preparation of p-iodonitrobenzene from p-nitroaniline.

5) Preparation of benzene azo - β - naphthol.

6) Preparation of benzoic acid from cinnamic acid.

IV Preparation of Derivatives :

N.B.:During practical course, name of the organic compound should not to be given.

- 1) Bromo derivative of aniline and cinnamic acid.
- 2) Nitro derivative of salicylic acid and nitrobenzene.
- 3) Benzoyl derivative of β -naphthol and aniline.
- 4) Picrate derivative of anthracene and β -naphthol.
- 5) Oxalate and nitro derivatives of urea.
- 6) Anhydride derivative of phthalic acid.
- 7) Oxime derivatives of Ketones : Acetone and acetophenone.
- 8) 2 : 4 DNP of acetophenone.

Reference Books :

1. Practical Organic Chemistry by A. I. Vogel.
 2. Hand book of Organic qualitative analysis by H. T. Clarke.
 3. A laboratory Hand Book of Organic qualitative analysis and separation by V. S. Kulkarni. Dastane Ramchandra & Co.
 4. Practical Organic Chemistry by F. G. Mann and B. C. Saunders. Low - priced Text Book. ELBS. Longman.
 5. Experiments in General Chemistry by C. N. R. Rao. Affiliated East-West Press Pvt. Ltd. Delhi.
 6. Advanced Practical Organic Chemistry by N. K. Vishnoi. Vikas Publishing House Private Limited.
 7. Comprehensive Practical Organic Chemistry Qualitative Analysis by V. K. Ahluwalia, Sunita Dhingra. University Press. Distributor - Orient Longman Ltd.
 8. Comprehensive Practical Organic Chemistry Preparation and Quantitative Analysis by V.K.Ahluwalia, Renu Aggarwal. University Press.Distributor-Orient Longmann Ltd.
 9. Practical Chemistry - Physical - Inorganic - Organic and Viva - voce by Balwant Rai Satija. Allied Publishers Private Limited.30
 10. College Practical Chemistry by H. N. Patel, S. R. Jakali, H. P. Subhedar, Miss. S. P. Turakhia. Himalaya Publishing House, Mumbai.
 11. College Practical Chemistry by Patel, Jakali, Mohandas, Israney Turakhia. Himalaya Publishing House, Mumbai.
 12. Practice of thin layer chromatography by Joseph C. Touchstone, Murrell F. Dobbins. A Wiley - Interscience Publication John-Wiley & Sons.
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